

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

General Certificate of Education O Level

MARK SCHEME for the June 2005 question paper

5070 CHEMISTRY

5070/02

Paper 2 (Theory 1), maximum mark 75

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JUNE 2005

GCE O Level

MARK SCHEME

MAXIMUM MARK: 75

SYLLABUS/COMPONENT: 5070/02

CHEMISTRY
Paper 2 (Theory 1)



Page 1	Mark Scheme	Syllabus	Paper
	O LEVEL – JUNE 2005	5070	2

Section A

Maximum 45 marks

- A1** four names at (1) each: penalise correct formulae once only
- (a) nitrogen dioxide
- (b) silicon dioxide
- (c) aluminium oxide
- (d) lead(II) iodide **[Total: 4]**
- A2** (a) iron has positive ions and delocalised electrons (1)
the electrons are free to move (1)
moving electrons is an electric current (1) **[3]**
- (b) high carbon steels are strong or are brittle (allow harder) (1)
low carbon steels are soft or are more easily shaped
(allow more malleable) (1) **[2]**
- (c) (i) conditions are air (oxygen) and water or moist air (1)
- (ii) magnesium is above iron in the reactivity series
(or is more reactive) (1)
hence it corrodes before the iron (1) **[3]**
- (d) any two from:
coloured compounds/variable oxidation states/can act as
catalysts/valency/form complex ions **[2]**
- (e) calculation
for idea of dividing by correct A_r (1)
dividing by the smallest (1)
for final formula only if first 2 fully correct (1)
- | | | | | | | | |
|---|----------|----|----------|---|----------|---|----------|
| K | 0.547/39 | Fe | 0.195/56 | C | 0.252/12 | N | 0.294/14 |
| | 0.0140 | | 0.00348 | | 0.0210 | | 0.0210 |
| | 4 | | 1 | | 6 | | 6 |
- i.e. $K_4FeC_6N_6$ or $K_4Fe(CN)_6$ **[3]**
- [Total: 13]**

Page 2	Mark Scheme	Syllabus	Paper
	O LEVEL – JUNE 2005	5070	2

- A3** (a) Group 0 or the noble gas group or Group 8 [1]
- (b) Any two sensible suggestions at (1) each
 e.g: Mendeleev's table has:
 Groups and periods reversed (only allow once)
 no A,
 no atomic numbers
 no transition metals
 periods 4 and/or 5 and all or a specific group has two elements
 group numbers Arabic rather than Roman [2]
- (c) any two observations at (1) each
 fizzes/runs on the surface/flame/dissolves/explodes/melts
 equation (1)
 $2 \text{Rb} + 2 \text{H}_2\text{O} \rightarrow 2 \text{RbOH} + \text{H}_2$ [3]
- [Total:6]**
- A4** (a) boiling point [1]
- (b) (i) making chemicals or feedstock or make petrol
not make plastics (1)
- (ii) for road surfaces (1) [2]
- (c) (i) saturated is single bonds or no double/triple bonds
or maximum number of hydrogen atoms (1)
 hydrocarbon is carbon and hydrogen only (1)
- (ii) correct methane structure (all dots = 1) (2) [4]
- (d) any two ideas at (1) each:
 enables supply to match demand (allow more useful)
 make more petrol
 make hydrogen
 make alkenes e.g. ethene [2]
- [Total: 9]**
- A5** (a) (i) hydrogen is below sodium in the reactivity series (1)
- (ii) chloride ions are removed (leaving hydroxide ions) (1) [2]
- (b) (i) chlorine bleaches litmus or turns starch/iodide paper blue (1)
- (ii) hydrogen pops with a burning splint (1) [2]
- (c) chlorine kills bacteria
 (not just sterilises the water) [1]
- (d) burning hydrogen does not produce pollutants or only forms
 water or hydrogen is not a finite resource, is renewable [1]

Page 3	Mark Scheme	Syllabus	Paper
	O LEVEL – JUNE 2005	5070	2

- (e) (i) no products or no reaction (1)
- (ii) sodium chloride and bromine, both needed for (1)
(allow NaCl and Br₂) [2]

[Total: 8]

- A6** (a) sodium ion shown as 2.8 (1)
chloride ion shown as 2.8.8 (1)
(charges not needed. Outer shell only = 0) [2]

- (b) (i) strong attraction between oppositely charged ions (1)
- (ii) higher charges on the ions (1)
hence stronger attraction (1)
(independent marks) [3]

- (c) ions cannot move in the solid but can move in the melt [1]

[Total: 6]

[Section A: score any 45 from 46]

Page 4	Mark Scheme	Syllabus	Paper
	O LEVEL – JUNE 2005	5070	2

Section B

Answer any three questions

- B7**
- (a) ozone is formed by photochemical reactions
(or sparks in air, u.v on O₂) [1]
- (b) ozone removed by reaction with chlorine (atoms) (1)
derived from CFC's (1)
ozone loss causes skin cancers or cataracts or crop damage
or skin diseases or eye damage (1)
(allow O₃ + CFC for (1)) [3]
- (c) (i) bond breaking is endothermic/absorbs energy (1)
and bond forming is exothermic/releases energy
more energy released than absorbed (only if first point scored) (1)
- (ii) as temperature increases molecules move faster
or increased k.e. (1)
hence more frequent collisions
or more molecules energy exceeds the activation energy (1)
- (iii) calculation
48 g ozone releases 143 kJ (1)
16 g ozone releases 47.66 kJ or 47.7 kJ (1)
(answer alone (1), units needed)
(if 6 x 16 = 96 g ozone used, then (0))
(if 0.33 used, answer = 47.2) [6]
- [Total: 10]
- B8**
- (a) calculation (2)
143.5 g AgCl contains 108 g Ag
0.287 g AgCl contains 0.216 g Ag
(answer alone (1), units needed) [2]
- (b) oxidation is electron loss or an increase in O.N. (1)
copper(I) is oxidised because it loses an electron
or its O.N. increases (1)
chlorine is reduced because it gains an electron
or its O.N. decreases (1) [3]
- (c) equation (1)
Ag + CuCl₂ → AgCl + CuCl [1]
- (d) (i) equation (1) state symbols (1)
CuCl₂(aq) + 2 NaOH(aq) → Cu(OH)₂(s) + 2 NaCl(aq),
(or ionic, Cu²⁺ + 2OH⁻ → Cu(OH)₂)
(scores (1) for states)
- (ii) name is copper(II) hydroxide (allow copper hydroxide) (1)
colour is blue or blue-green (1)
(colour only for correct name) [4]
- [Total: 10]

Page 5	Mark Scheme	Syllabus	Paper
	O LEVEL – JUNE 2005	5070	2

- B9** (a) (i) the catalyst is iron or Fe_2O_3 (1)
- (ii) equation
 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ (1)
- (iii) the temperature is 280°C (1)
the pressure is 400 atmos (1)
- (iv) higher temperature gives faster reaction (1)
(higher yield = -1) [5]
- (b) a catalyst increases reaction rate (1)
(not alters the rate)
a lower activation energy (1)
hence saves energy (1)
(third mark only if E_a given) [3]
- (c) equation (1)
 $\text{Ca}(\text{OH})_2 + 2\text{NH}_4\text{NO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + 2\text{H}_2\text{O} + 2\text{NH}_3$
ammonia lost as a gas (1) [2]
- [Total: 10]
- B10** (a) name is butanoic acid (not butenoic) (1)
- (b) formula is $\text{C}_5\text{H}_{11}\text{CO}_2\text{H}$ (not $\text{C}_6\text{H}_{12}\text{O}_2$) (1)
- (c) structure of ethyl ethanoate (1)
allow full structure or condensed version, $\text{CH}_3\text{CO}_2\text{C}_2\text{H}_5$
- (d) allow any suitable named oxidising reagent (1)
e.g. (acidified) potassium dichromate(VI) or air or oxygen
(allow formula) [(a) to (d) 4]
- (e) equation (1)
 $\text{Mg} + 2\text{CH}_3\text{CO}_2\text{H} \rightarrow \text{Mg}(\text{CH}_3\text{CO}_2)_2 + \text{H}_2$
calculation (2)
 50 cm^3 acid is 0.05 mol 0.025 mol Mg needed
 $24 \times 0.025 = 0.60\text{ g}$
(answer alone (1), unit needed) [3]
- (f) ethanoic acid is weak and hydrochloric is strong (1)
lower $[\text{H}^+]$ concentration in ethanoic acid (1) [2]
- (g) ionic equation (1)
 $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ [1]
- [Total: 10]